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thermodynamics. chapter 04: entropy and the

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Apr 23, 2018 · Based on graph below, what is the work done in process A-B-C-A. Solution :
Work (W) = Area of the triangle A-B-C. $W = \frac{1}{2} (20-10) (6 \times 10^5 - 2 \times 10^5)$ $W = \frac{1}{2} (10) (4 \times 10^5)$ $W = (5) (4 \times 10^5)$ $W = 20 \times 10^5$. $W = 2 \times 10^6$ Joule. Read : Dimensional analysis – problems and solutions.

Given that the free energy of formation of liquid water is -237 kJ / mol , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE^\circ$ and $n = 4$, we calculate the

potential is -1.23 V.

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ft g =75 g =5.4 g =32 sec sec lbf-sec² c moon
cmoon Jupiter Jupiter c mg Wg10×32 W = m
= = 59.26 lb gg5.4 mg 59.26×75 W = 139 lbf
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Problem 1:-A container holds a mixture of three nonreacting gases: n_1 moles of the first gas with molar specific heat at constant volume C_{v1} , and so on. Find the molar specific heat at constant volume of the mixture, in terms of the molar specific heats and quantities of the three separate gases.

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Solutions to extra problems in Chapter 11:
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runs a Rankine cycle with a water boiler at 3.5
MPa and the cycle has ...
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Answers For Thermodynamics Problems.
Answer for Problem # 1. Since the containers
are insulated, no heat transfer occurs between
the gas and the external environment, and
since the gas expands freely into container B
there is no resistance "pushing" against it,

which means no ...

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